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Chapter 11 Review Molecular Composition

CHAPTER 11 REVIEW Molecular Composition of Gases SECTION 11-3 SHORT ANSWER Answer the following questions in the space provided. 1. Volumes of gaseous reactants and products in a chemical reaction can be expressed as ratios of small whole numbers if . (a) all reactants and products are gases (b) standard temperature and pressure are maintained

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REVIEW. SHORT ANSWER Answer the following questions in the space provided. 1. c The average speed of a gas molecule is most directly related to the . (a) polarity of the molecule. (b) pressure of the gas. (c) temperature of the gas. (d) number of moles in the sample.

11 Molecular Composition of Gases - Madison Public Schools

Molecular Composition of Gases Chapter 11. What restrictions are there on the use.... At the same temperature and pressure, w.... According to Avogadro, what is the rela.... What is the relationship between the nu.... It applies only to gases measured under the same temperature a....

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Chapter 11 - Molecular Composition of Gases 11-1 Volume-Mass Relationships of Gases I. Measuring and Comparing the Volumes

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of Reacting Gases A. Observations of Gay-Lussac 1. 2 liters H₂ + 1 liter O₂ \rightarrow 2 liters H₂O vapor 2. 2 volumes H₂ + 1 volume O₂ \rightarrow 2 volumes H₂O vapor 3. 1 volume H₂ + 1 volume Cl₂ \rightarrow 2 volumes HCl 4.

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REVIEW SHORT ANSWER Answer the following questions in the space provided
1 c The average speed of a gas molecule is most directly related to the (a) polarity of the molecule (b) pressure of the gas (c) temperature of the gas (d) number of moles in the sample
2 d Which of

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[MOBI] Chapter 11 Review Gases Section 2

Chapter 11 Review TRUE/FALSE. Write 'T' if the statement is true and 'F' if the statement is false. 1) Light from distant stars that must pass through dust arrives bluer than when it left its star. 1) 2) Interstellar matter is distributed very evenly throughout the galaxy. 2) 3) The gas in the interstellar medium consists of mostly heavy ...

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334 CHAPTER 11 FIGURE 11-1 At the same temperature and pressure, balloons of equal volume have equal numbers of molecules, regardless of which gas they contain. Hydrogen molecule 1 mol H₂ at STP = 22.4 L Oxygen molecule 1 mol O₂ at STP = 22.4 L Carbon dioxide molecule 1 mol CO₂ at STP = 22.4 L

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Chapter 11 Review Gases Section 2 Answers - Bing | pdf ... explains some of the properties of ideal gases. In this chapter, you will study the predictions of kinetic-molecular theory for gases in more detail. This includes the relationship among the temperature, pressure, volume, and amount of gas in a sample.

Chapter 11 Review Gases Section 1 Answers

Answer Key Chapter 11 - Chemistry 2e | OpenStax 1. A solution can vary in composition, while a compound cannot vary in composition. Solutions are homogeneous at the molecular level, while other mixtures are heterogeneous.

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